Science 10 – Chapter 5.2 - Salts Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_

**Acid-Base Neutralization**

Salts are a class of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ that can be formed during the

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. They are composed of a

positive ion (\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_) and a negative ion (\_\_\_\_\_\_\_\_\_\_\_\_\_\_)

Neutralization reactions occur when an\_\_\_\_\_\_\_\_ and a \_\_\_\_\_\_\_\_ react to produce a \_\_\_\_\_\_\_\_\_\_ and

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Eg. HCl(aq) + NaOH(aq) →

*acid base salt water*

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**Metal Oxides and Non-Metal Oxides**

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ react with water to form \_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Ex. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + H2O(l ) → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(aq)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ react with water to form \_\_\_\_\_\_\_\_\_\_\_\_\_.

Ex. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + H2O(l ) → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(aq)

Non-metal oxides are formed from the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Add water in the atmosphere = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Acids and Metals**

The most reactive metals, at the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, react

vigorously with \_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

All other metals are \_\_\_\_\_\_\_\_\_ reactive than those in groups 1 and 2.

When metals do react with acids, \_\_\_\_\_\_\_\_\_\_ gas is usually released and a \_\_\_\_\_\_\_\_\_\_\_ is produced.

Ex. \_\_\_\_\_\_\_\_\_\_\_\_\_\_(aq) + \_\_\_\_\_\_\_\_ (s) → \_\_\_\_\_\_\_\_\_\_\_ (s) + \_\_\_\_\_\_\_\_\_\_\_(g)

Do practice problems on p238

**Acids and Carbonates**

Carbonates (-CO3) \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ acids, protecting locations with natural carbonate supplies from

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Ex. H2SO4 + CaCO3 → \_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_

(acid) (carbonate) (salt) (water) (carbon dioxide)

Adding lime (calcium carbonate - CaCO3) to a lake is called “liming.”



When and why would this be helpful?

Why is it not practical to always use liming?