1) \_\_\_ CuI2 + \_\_\_ Fe 🡪 \_\_\_ FeI2 + \_\_\_ Cu

**Balancing Equations Practice**  Date: Name:

2) \_\_\_ Sn + \_\_\_ O2 🡪 \_\_\_ SnO

3) \_\_\_ Fe + \_\_\_ Br2 🡪 \_\_\_ FeBr3

4) \_\_\_ V2O5 + \_\_\_ Ca 🡪 \_\_\_ CaO + \_\_\_ V

5) \_\_\_ C2H6 + \_\_\_ O2 🡪 \_\_\_ CO2 + \_\_\_ H2O

6) \_\_\_ NH3 + \_\_\_ O2 🡪 \_\_\_ N2 + \_\_\_ H2O

7) \_\_\_ Pb(NO3)2 + \_\_\_ K2CrO4 🡪 \_\_\_ PbCrO4 + \_\_\_ KNO3

8) \_\_\_ Sn(NO2)4 + \_\_\_ K3PO4 🡪 \_\_\_ KNO2 + \_\_\_ Sn3(PO4)4

9) sodium + water 🡪 sodium hydroxide + hydrogen

10) iron (III) oxide + hydrogen 🡪 water + iron

11) chromium + tin (IV) chloride 🡪 chromium (III) chloride + tin

10) Chromium (III) sulphate reacts with potassium carbonate to form chromium (III) carbonate and potassium sulphate.

**Be careful of diatomic elements!**

The “special seven” are all diatomic elements (hockey stick & puck).

They are: H2, N2, O2, F2, Cl2, Br2, I2

