

Name: _____

Block: _____

Date: _____

Science 10

4.1 Bohr Models & Lewis Structures

1. Complete the following Bohr & Lewis diagrams.

$\frac{5}{\text{B}}$ Boron $\frac{10.8}{}$	P = <u>5</u> N = <u>2</u> E = <u>5</u>
Bohr Diagram	
Lewis Structure	
$\cdot \overset{\cdot}{\underset{\cdot}{\text{B}}}$	

$\frac{3}{\text{Li}}$ Lithium $\frac{6.9}{}$	P = <u>3</u> N = <u>2</u> E = <u>3</u>
Bohr Diagram	
Lewis Structure	
$\overset{\cdot}{\text{Li}}$	

$\frac{10}{\text{Ne}}$ Neon $\frac{20.2}{}$	P = <u>10</u> N = <u>10</u> E = <u>10</u>
Bohr Diagram	
Lewis Structure	
$:\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Ne}}}$	

$\frac{2}{\text{He}}$ Helium $\frac{4.0}{}$	P = <u>2</u> N = <u>2</u> E = <u>2</u>
Bohr Diagram	
Lewis Structure	
$\overset{\cdot\cdot}{\text{He}}$	

$\frac{6}{\text{C}}$ Carbon $\frac{12.0}{}$	P = <u>6</u> N = <u>6</u> E = <u>6</u>
Bohr Diagram	
Lewis Structure	
$\cdot \overset{\cdot}{\underset{\cdot}{\text{C}}}$	

$\frac{15}{\text{P}}$ Phosphorus $\frac{31.0}{}$	P = <u>15</u> N = <u>16</u> E = <u>15</u>
Bohr Diagram	
Lewis Structure	
$\cdot \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}}$	

$\frac{16}{\text{S}}$ Sulfur $\frac{32.1}{}$	P = <u>16</u> N = <u>16</u> E = <u>16</u>
Bohr Diagram	
Lewis Structure	
$\cdot \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{S}}}$	

$\frac{12}{\text{Mg}}$ Magnesium $\frac{24.3}{}$	P = <u>12</u> N = <u>12</u> E = <u>12</u>
Bohr Diagram	
Lewis Structure	
$\overset{\cdot}{\text{Mg}}$	

$\frac{1}{\text{H}}$ Hydrogen $\frac{1.0}{}$	P = <u>1</u> N = <u>0</u> E = <u>1</u>
Bohr Diagram	
Lewis Structure	
$\overset{\cdot}{\text{H}}$	