

Chapter 4.2 & 4.3 Practice Test

Name:

- Which one of the following elements is multivalent?
 - lithium
 - fluorine
 - nitrogen
 - nickel
- What is the name of the ionic compound, BeBr_2 ?
 - beryllium bromine
 - beryllium bromide
 - beryllium dibromide
 - monoberyllium dibromide
- What is the name of the covalent compound, SiO_2 ?
 - silicon dioxide
 - silicon (II) oxide
 - monosilicon dioxide
 - silicon oxide (II)
- What is the name of the ionic compound, K_3P ?
 - tripotassium phosphide
 - potassium (III) phosphide
 - potassium phosphate
 - potassium phosphide
- What is the name of the covalent compound, N_2O_3 ?
 - nitrogen oxide
 - nitrous oxide
 - dinitrogen trioxide
 - nitrogen (II) trioxide
- What is the name of the compound, PbCl_2 ?
 - lead chloride
 - lead (II) chloride
 - lead (IV) chloride
 - lead dichloride
- What is the name of the compound, NaNO_2 ?
 - sodium nitrite
 - sodium nitrate
 - sodium nitrogen oxide
 - sodium nitrogen dioxide
- What is the formula of the compound, diphosphorous trisulphide?
 - PS
 - PS_3
 - P_3S_2
 - P_2S_3
- What is the formula for the compound, manganese (III) oxide?
 - MnO
 - Mn_3O
 - MnO_2
 - Mn_2O_3
- What is the formula for the compound, potassium phosphite?
 - K_3PO_3
 - K_3PO_4
 - KP
 - $\text{K}(\text{PO}_3)_3$
- Why is the name monocarbon monoxide incorrect for the formula CO ?
 - CO is an ionic compound which means no prefixes when naming
 - Mono is never used before the first name
 - Monocarbon should be written as monocarbide
 - Monoxide should be written as monoxide
 - I and III
 - I and IV
 - II and III
 - II and IV
- Which of the following correctly classifies each formula as an ionic compound, a polyatomic compound or a covalent compound?

	Ionic	Polyatomic	covalent
a.	CuCl_2	KNO_3	CCl_4
b.	CCl_4	CuCl_2	KNO_3
c.	KNO_3	CCl_4	CuCl_2
d.	CuCl_2	CCl_4	KNO_3
- Which set of ordered coefficients correctly balances the following equation?

$$\underline{\hspace{1cm}} \text{K}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{ZnSO}_4 \rightarrow \underline{\hspace{1cm}} \text{K}_2\text{SO}_4 + \underline{\hspace{1cm}} \text{Zn}_3(\text{PO}_4)_2$$

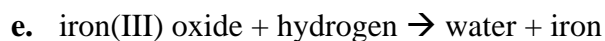
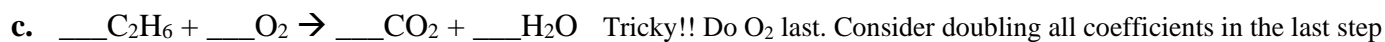
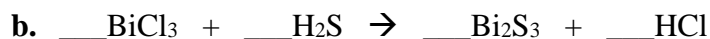
 - 1, 2, 3, 2
 - 2, 3, 3, 1
 - 2, 1, 3, 2
 - 2, 2, 1, 3

PART 2: WRITTEN RESPONSE (15 marks)

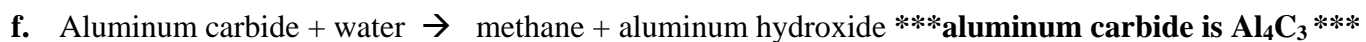
1. Please complete the following table. Hint: you will need to use your data sheet to complete the formulas for the ionic and polyatomic compounds. (7 marks)

Scientific Name	Formula
silver phosphide	
	NH_4NO_2
manganese (II) sulphide	
	CuCl_2
nitrogen tribromide	
lead (II) perchlorate	
	Cl_2O

2. Balance the following equations (8 marks)



(2 marks – 1 for correctly writing the equation, 1 for balancing)



(2 marks – 1 for correctly writing the equation, 1 for balancing)